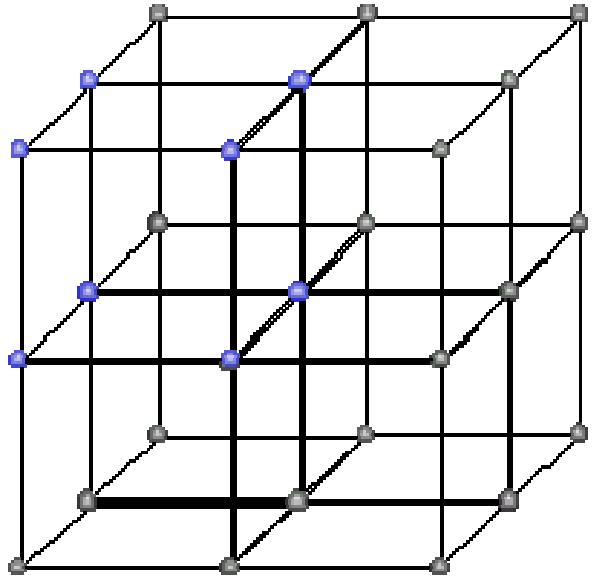
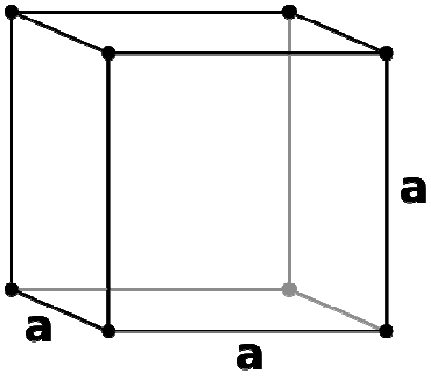
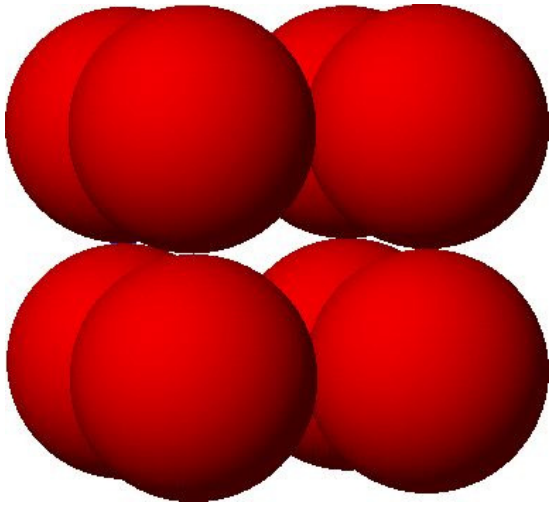
Slides
Condensed Matter Physics
Lecture 10

Simple (Primitive) Cubic Lattice

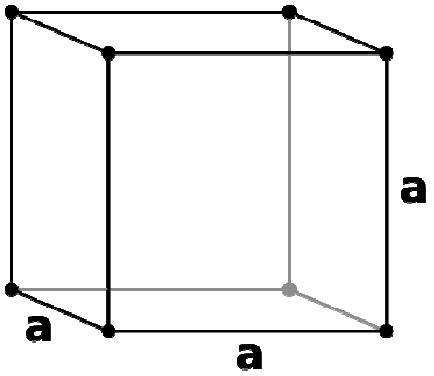


Simple Cubic Unit Cell



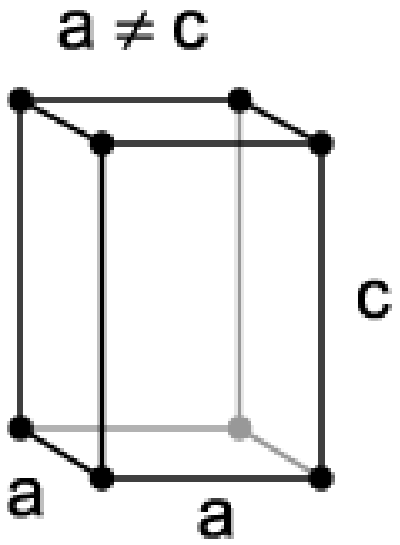


Atoms arranged in Simple Cubic Lattice (very unusual)

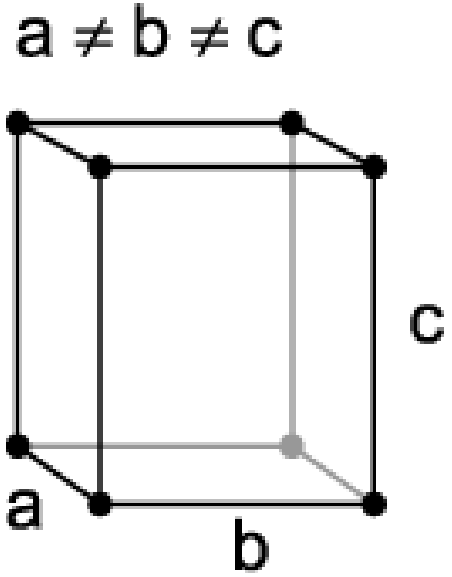


Simple Cubic Unit Cell

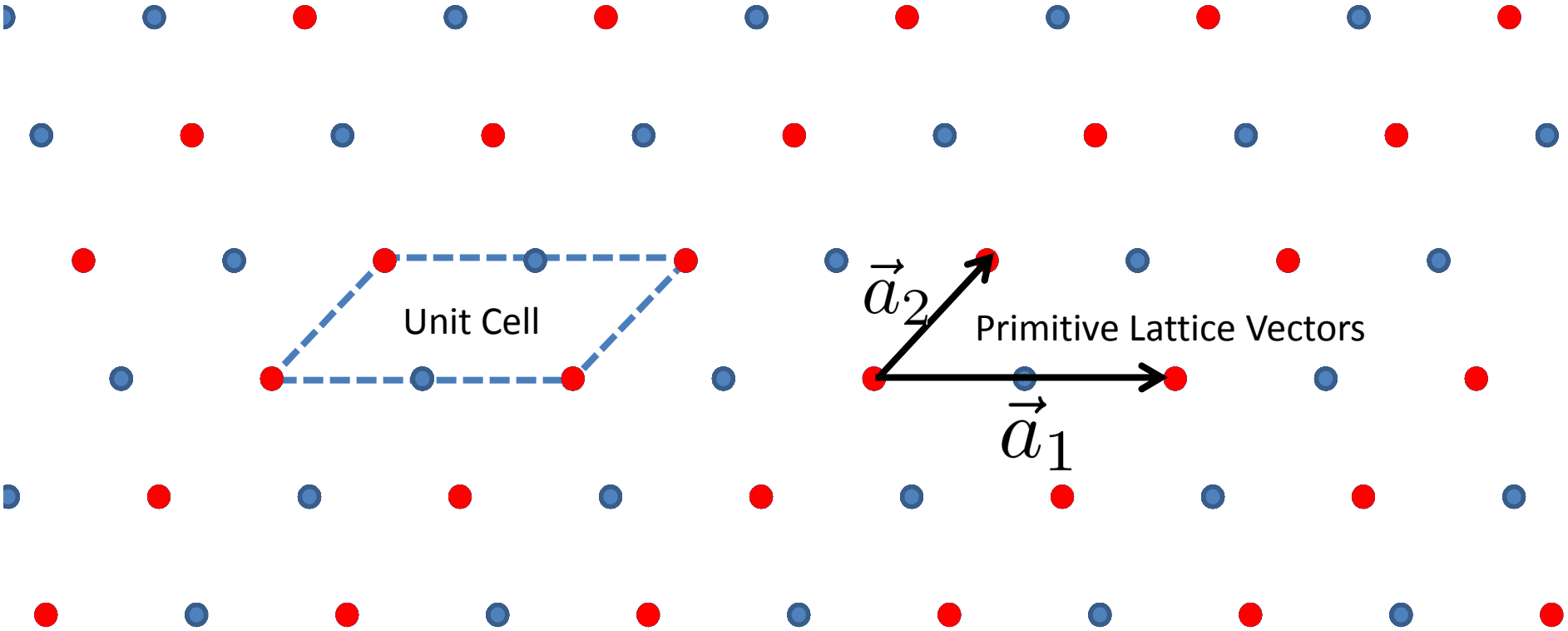
Simple Tetragonal Unit Cell



Simple Orthorhombic Unit Cell



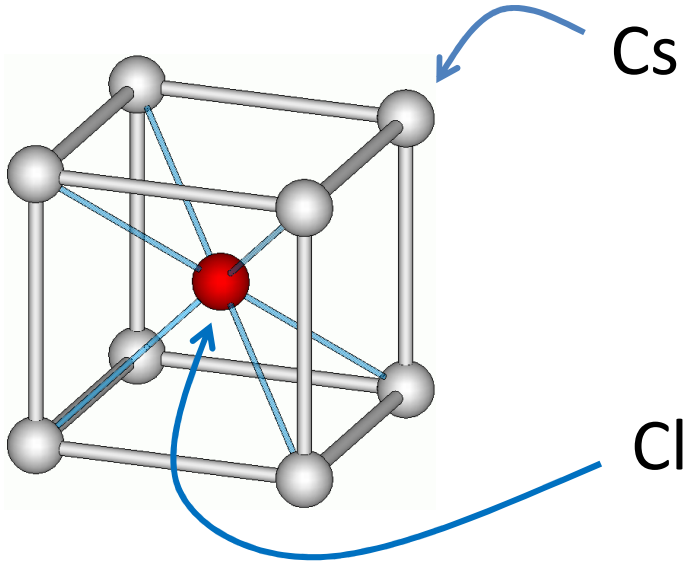
Reference Lattice is often taken coincident with some atom



Put Reference Lattice on the Red Atoms:

Basis is: Red atom at [0,0]
Blue atom at [1/2,0]

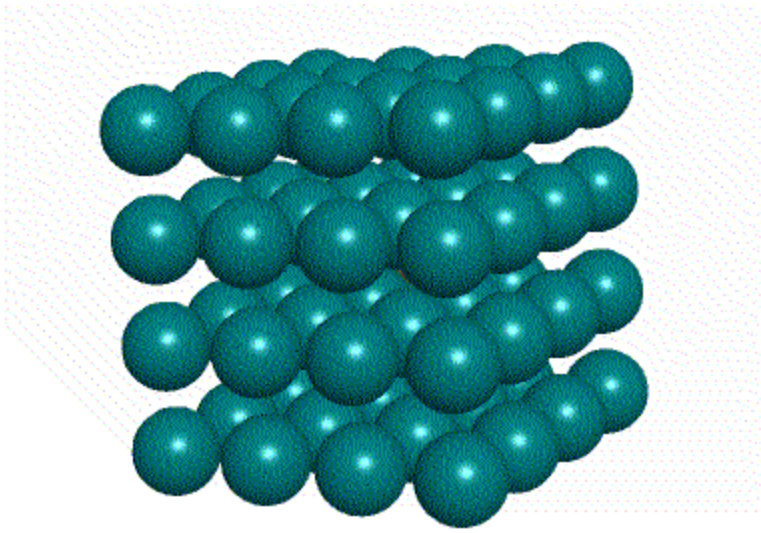
note [1/2,0] = $(1/2)\vec{a}_1$

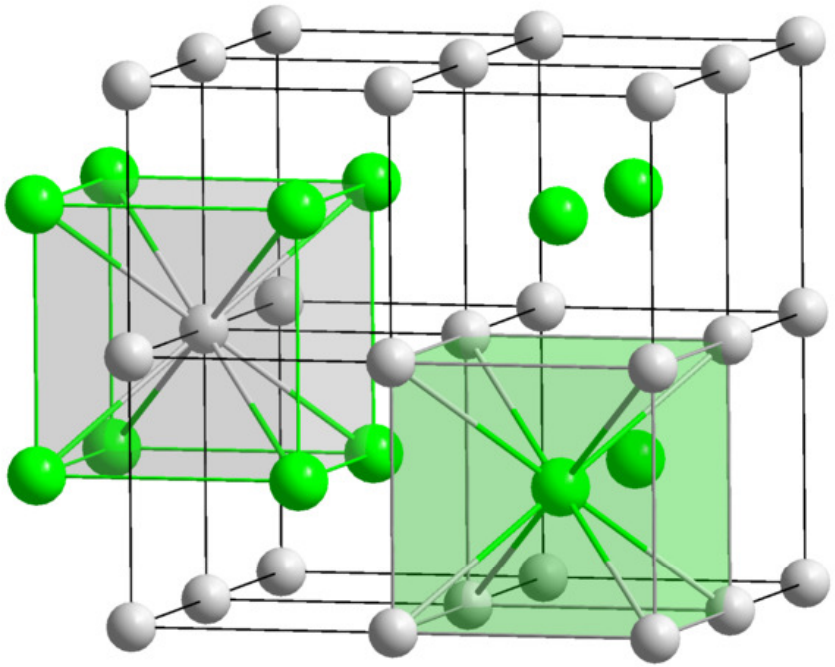


Cesium Chloride (CsCl): A simple cubic Lattice with a Basis

Basis:

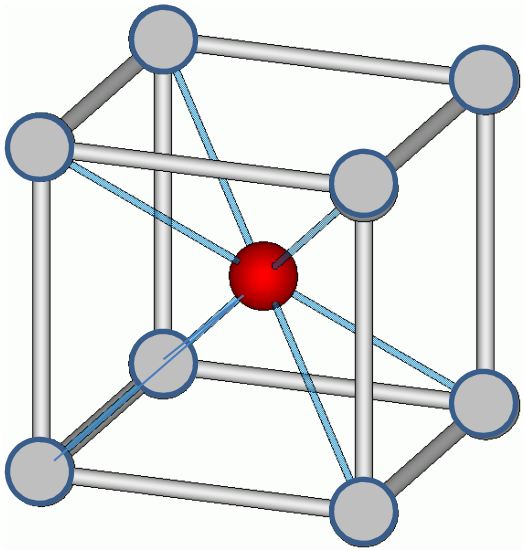
Cs at[0, 0, 0] (i.e., on the simple cubic)
Cl at [½, ½, ½] (i.e., in the middle of each cube)





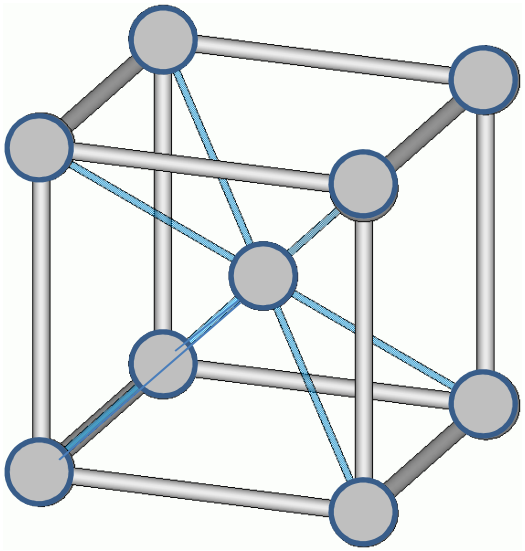
Two depictions of the CsCl lattice structure =

Two interlocking simple cubics



CsCl =Simple Cubic
with Basis

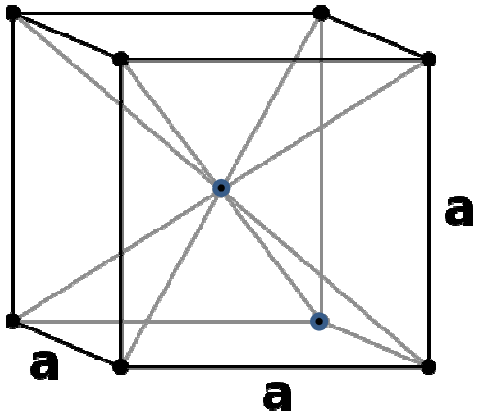
Cs at [0, 0, 0]
Cl at [½, ½, ½]



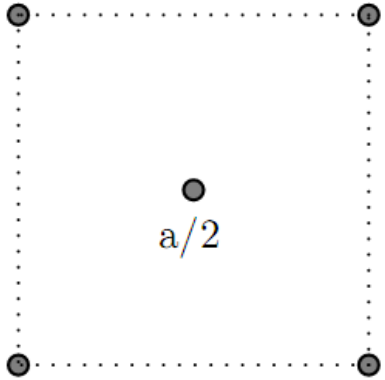
Cs = Simple Cubic
with Basis

Cs at [0, 0, 0]
Cs at [½, ½, ½]

Unit cell of Body Centered Cubic Lattice (BCC) (Notated cubic-I)

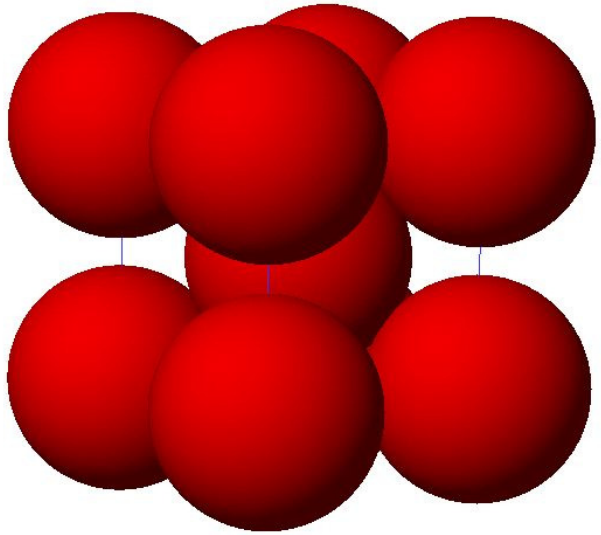


Conventional Unit Cell

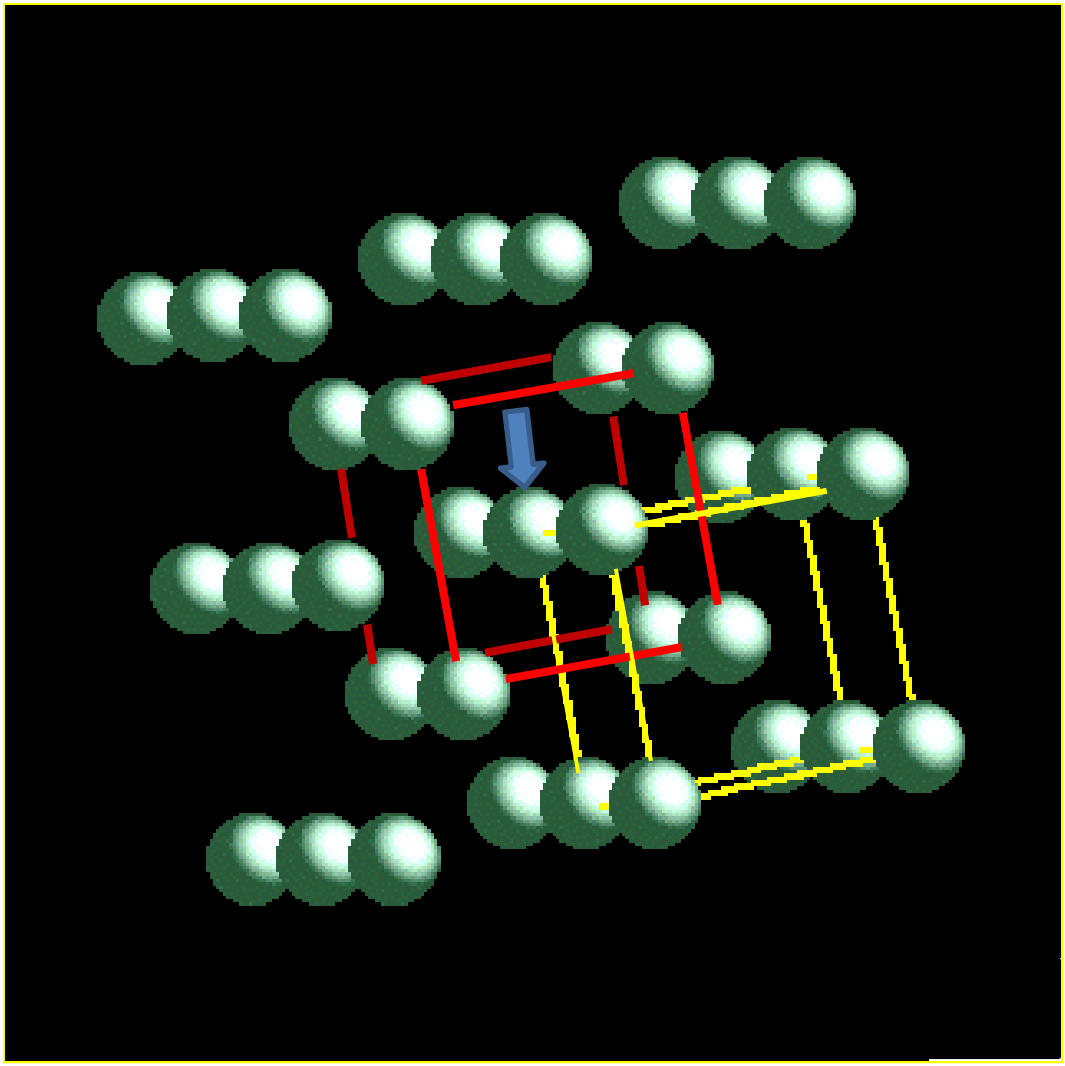


Plan View

(unlabeled points at height 0 and a)

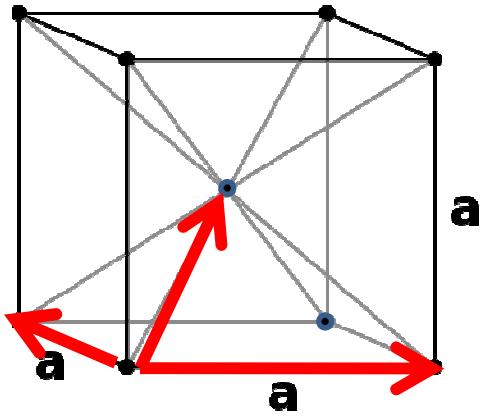


(More efficient sphere packing)

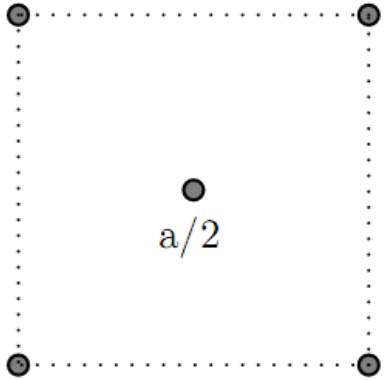


BCC Lattice

Unit cell of Body Centered Cubic Lattice (BCC) (Notated cubic-I)

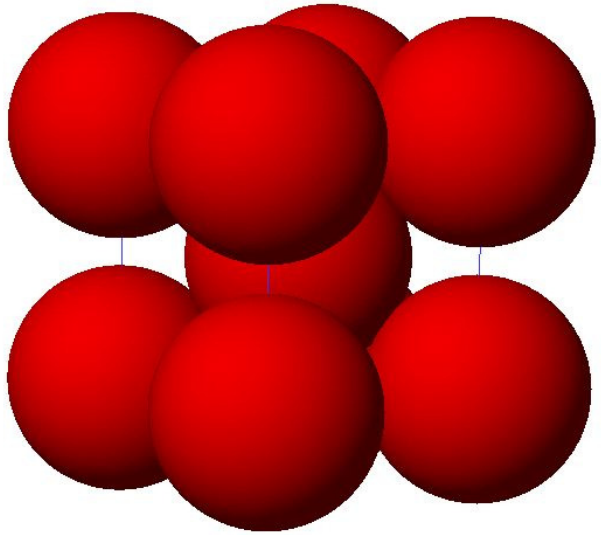


Conventional Unit Cell

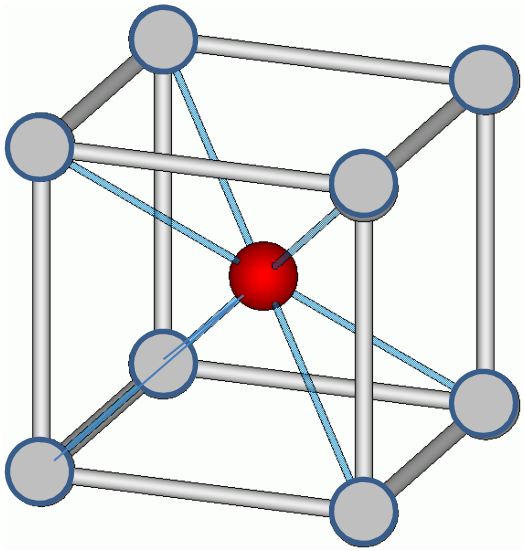


Plan View

(unlabeled points at height 0 and a)

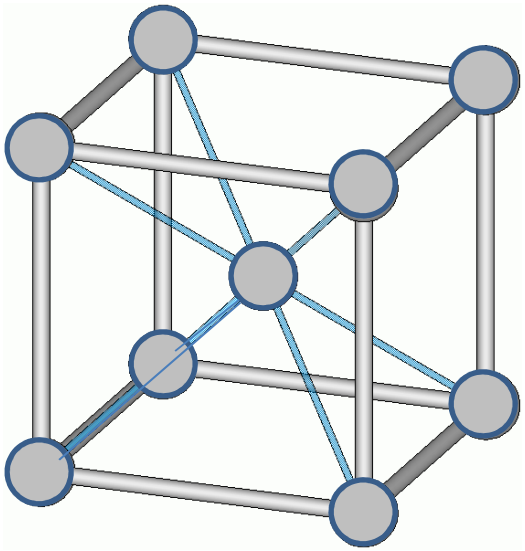


(More efficient sphere packing)



CsCl =Simple Cubic with Basis

Cs at [0, 0, 0]
Cl at [½, ½, ½]



Cs = Simple Cubic with Basis

Cs at [0, 0, 0]
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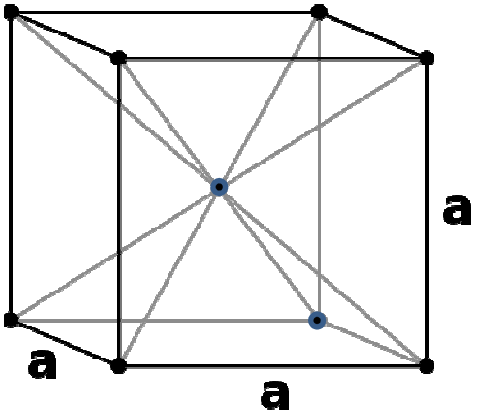
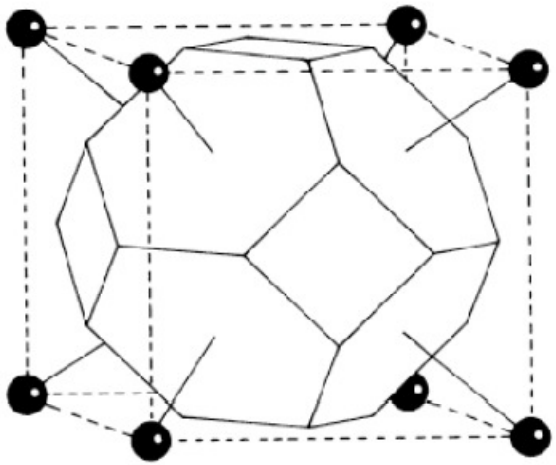
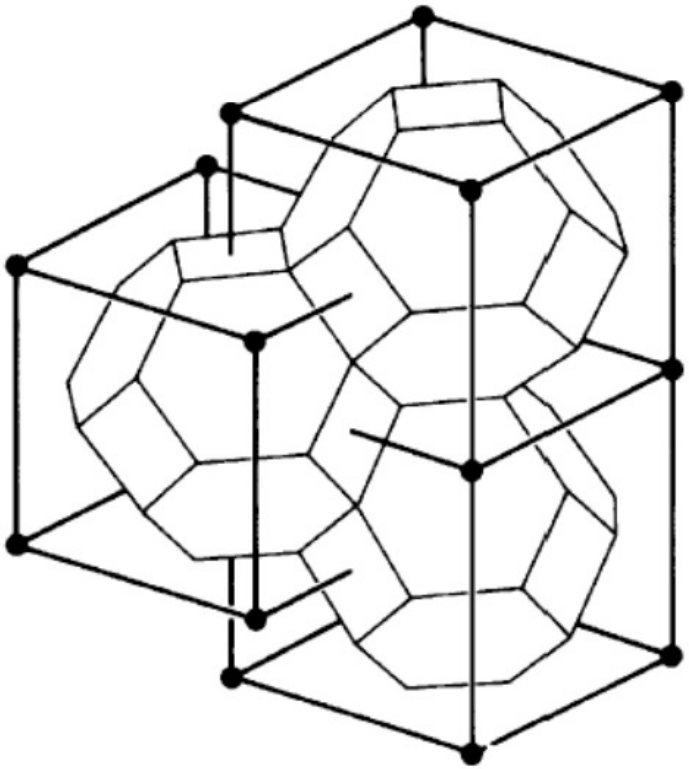
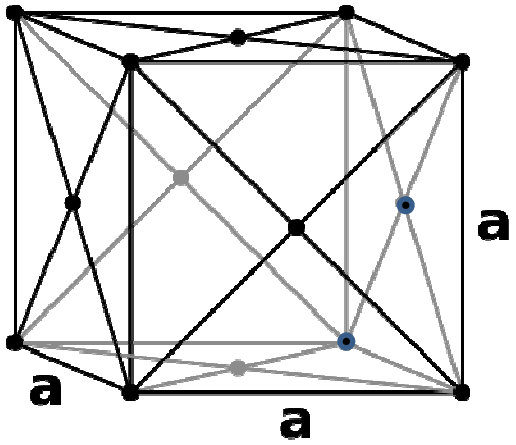




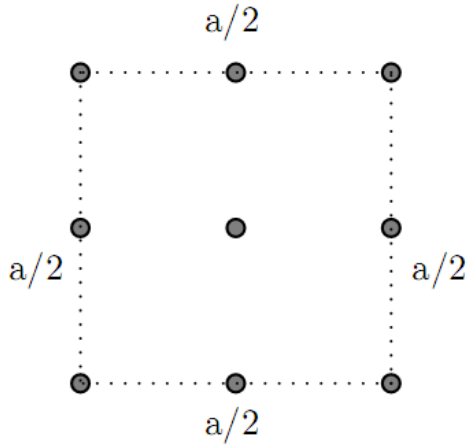
The Wigner-Seitz (Primitive) Unit Cell for the BCC lattice



Unit cell of Face Centered Cubic Lattice (FCC) (Notated cubic-F)

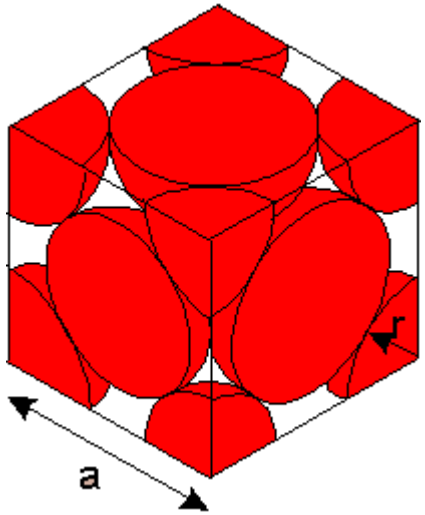


Conventional Unit Cell



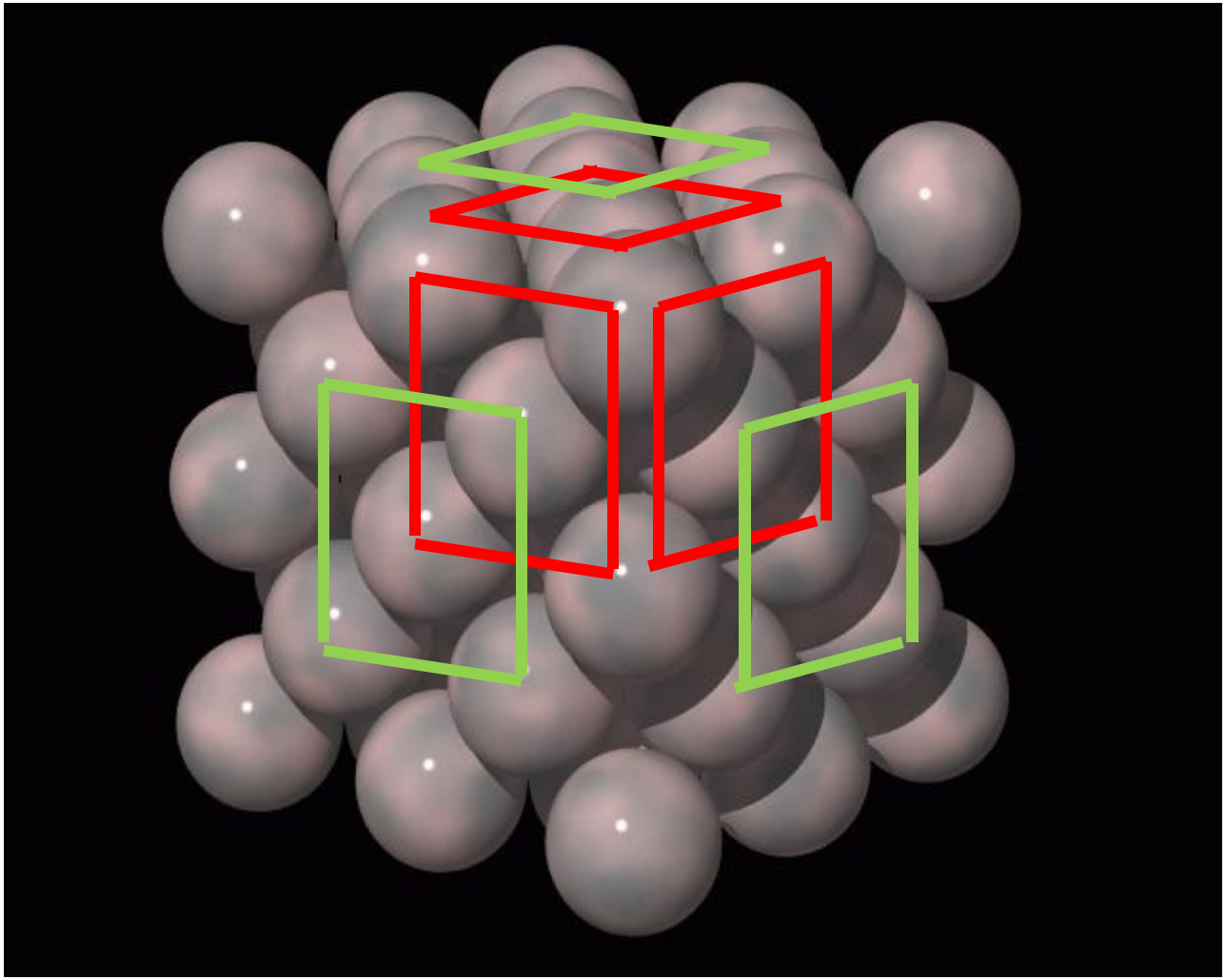
Plan View

(unlabeled points at height 0 and a)

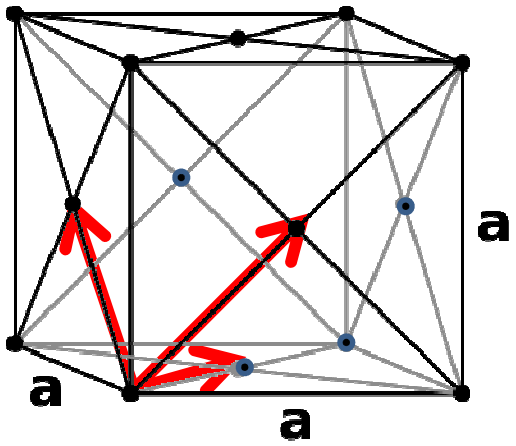


Most efficient sphere packing

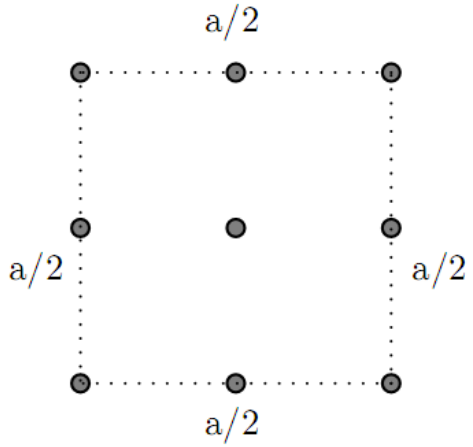
same density as “hcp” packing which we don’t’ learn this year



Unit cell of Face Centered Cubic Lattice (FCC) (Notated cubic-F)



Conventional Unit Cell



Plan View

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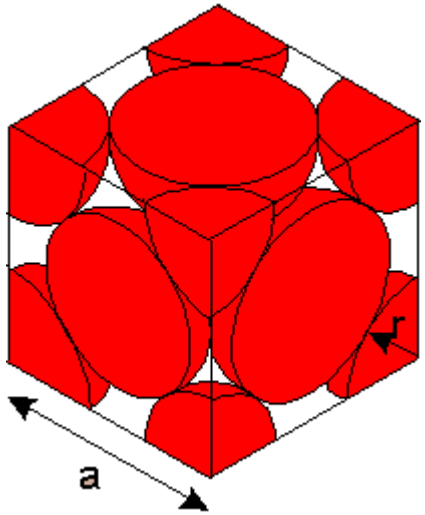
FCC lattice



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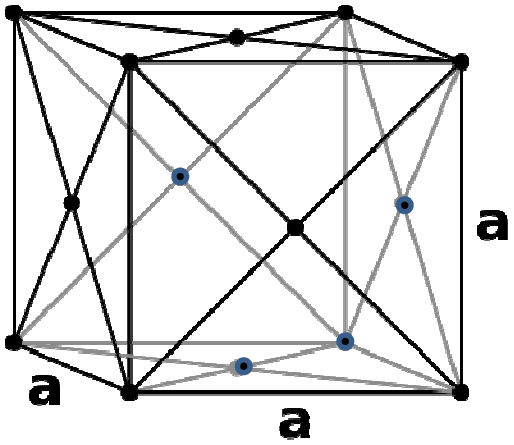
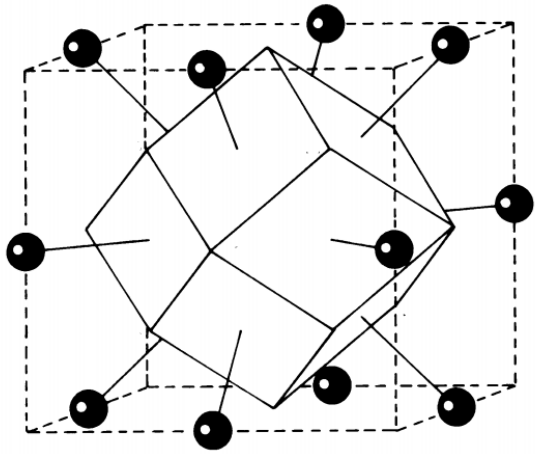


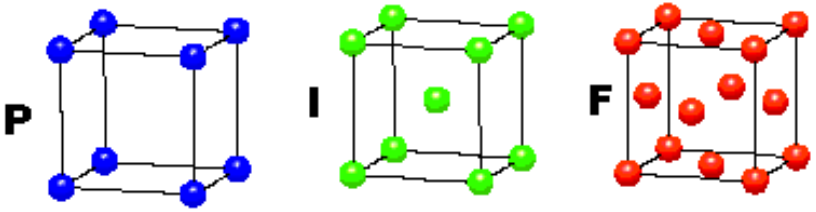


The Wigner-Seitz (Primitive) Unit Cell for the FCC lattice

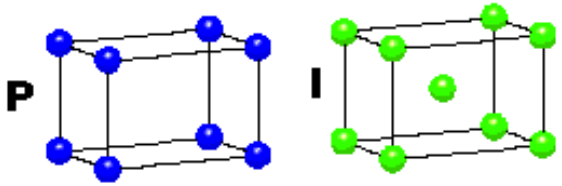
Packing Wigner Seitz cells to fill space

The 14 Bravais Lattice Types

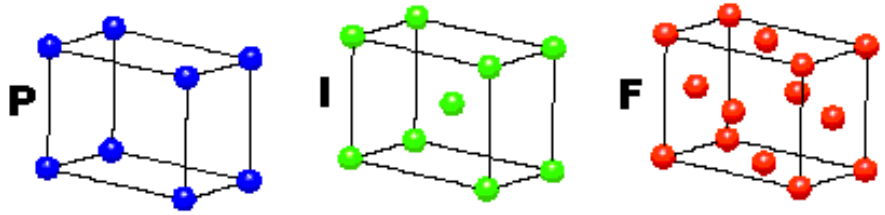
CUBIC
$a = b = c$
$\alpha = \beta = \gamma = 90°$



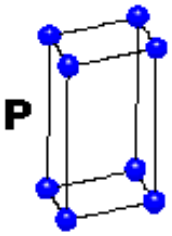
TETRAGONAL
$a = b \neq c$
$\alpha = \beta = \gamma = 90°$



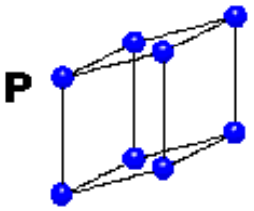
ORTHORHOMBIC
$a \neq b \neq c$
$\alpha = \beta = \gamma = 90°$



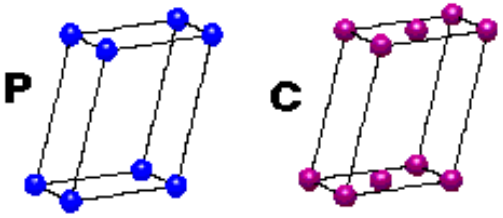
HEXAGONAL
$a = b \neq c$
$\alpha = \beta = 90°$
$\gamma = 120°$



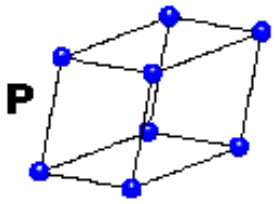
TRIGONAL
$a = b = c$
$\alpha = \beta = \gamma \neq 90°$



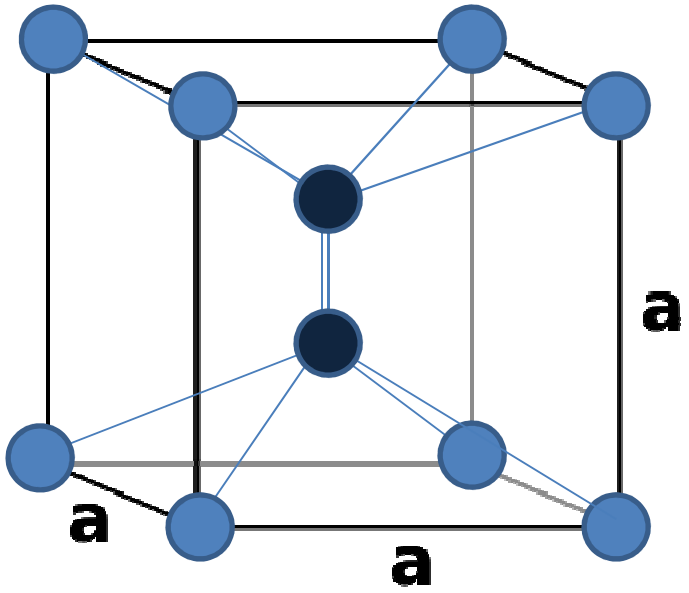
MONOCLINIC
$a \neq b \neq c$
$\alpha = \gamma = 90°$
$\beta \neq 120°$



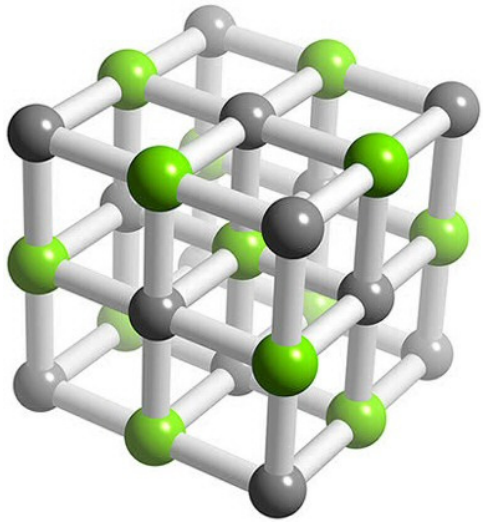
TRICLINIC
$a \neq b \neq c$
$\alpha \neq \beta \neq \gamma \neq 90°$



4 Types of Unit Cell
P = Primitive
I = Body-Centred
F = Face-Centred
C = Side-Centred
+
7 Crystal Classes
→ 14 Bravais Lattices



Subtlety: This is NOT cubic
(does not look the same from six sides)



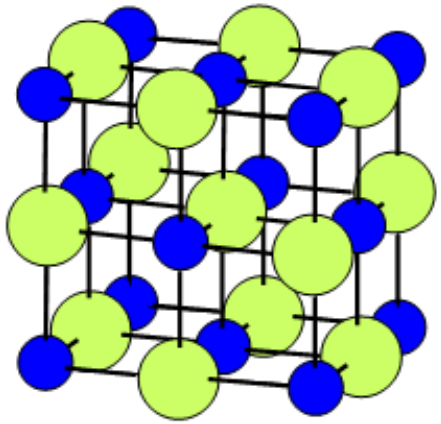
sodium chloride (NaCl)

lattice: cubic F

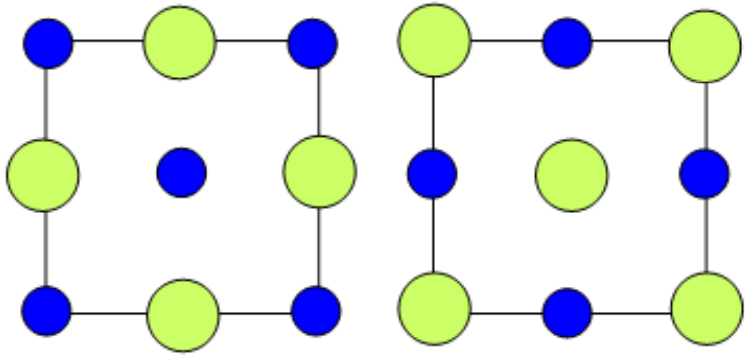
basis :

Na 000

Cl ½½½



Plan view



$z = 0$ layer $z = ½$ layer

Na forms FCC lattice.

Cl is displaced [½, ½, ½] from each Na

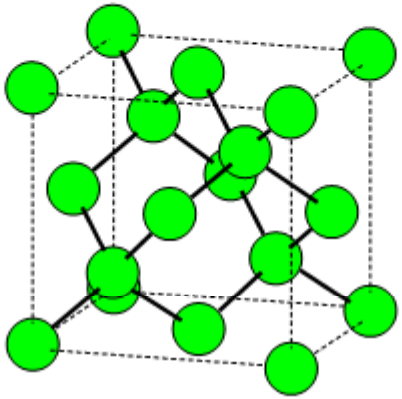
diamond (C) — also Si, Ge

lattice: cubic F

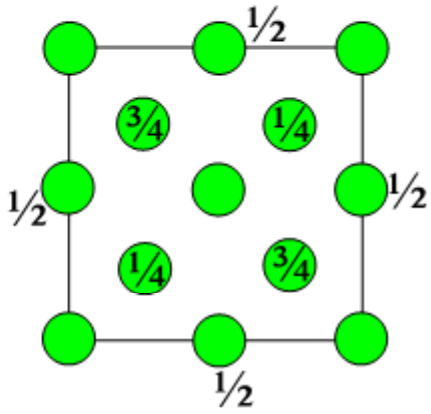
basis :

C 000

C ¼¼¼

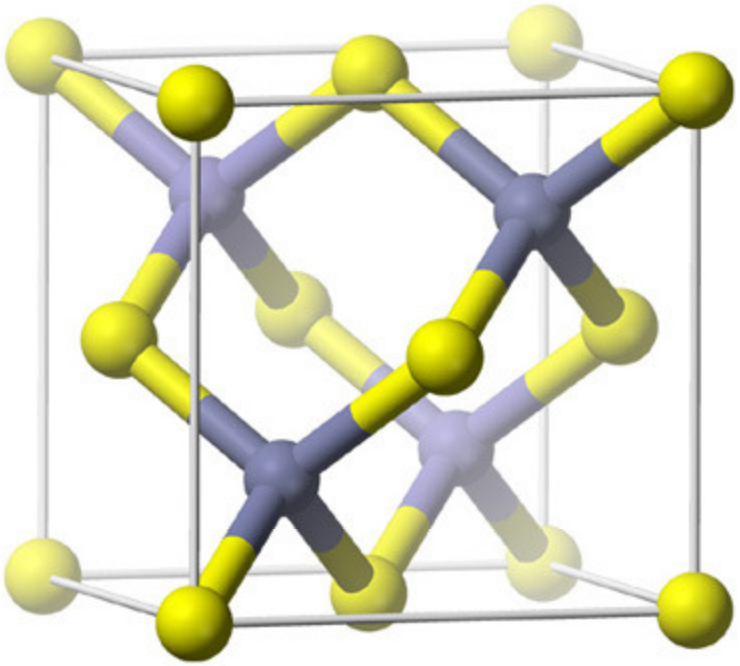


Plan view

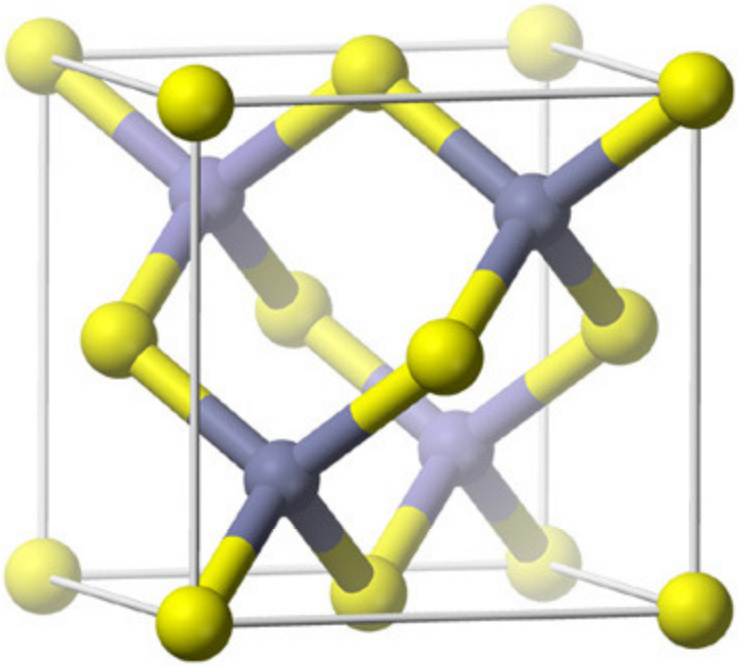


(unlabled points at height 0 and 1)

GaAs Structure…



GaAs Structure…



FCC:

Basis: Yellow at [0, 0, 0]

Blue at [¼ , ¼, ¼]